

V Semester B.Sc. Examination, Nov./Dec. 2018 (CBCS) (F+R) (2016-17 and Onwards) CHEMISTRY - V **Organic Chemistry**

Time: 3 Hours

Max. Marks: 70

Instructions: 1) The question paper has two parts. Answer both the parts.

2) Draw diagrams and chemical equations wherever necessary.

PART - A

Answer any eight of the following questions. Each question carries

- 1. Write R and S configuration of 2-amino propanoic acid.
- 2. What are the necessary conditions for a cyclic compounds to exhibit geometrical isomerism?
 - 3. How is acetone converted into isopropyl amine?
 - 4. Methanamine is more basic than ammonia, give reason
 - 5. How is pyrrole prepared from ammonium mucate?
 - 6. Mention one medicinal use of :
 - i) Ephedrine
 - ii) Caffeine.
 - 7. State "Isoprene rule".
 - 8. Why TMS uses as reference in NMR spectroscope?
 - 9. Mention the electronic transition that takes place when UV radiation is passed through acetone.

18. a) Discuss the general mechanism of electrophilic substitution of pyrrole.

b) What happens when quinoline is nitrated? Give equation.

(4+2)

(4+2)

Give the reaction.



19.	a)	Write general characteristics of alkaloids.	
9	b)	How would you prove that α -terpineol contains one double bond ?	4+2)
20.	a)	How is glucose converted to fructose?	
			(4+2)
21	a)	Describe the synthesis of nicotine from succinamide.	
			(4+2)
		How do you prove the presence of six membered ring in glucose by	
22		periodic acid oxidation method ?	
	b)	Write the structure of dichlofenac and mention its use BMSCA	(4+2)
000		Explain the effect of conjugation on the UV spectra of organic compound	S
23). a)	with an example.	
-	b	How is IR spectrum used to distinguish between free -OH and hydrogen	
		bonded -OH groups ?	(4+2)
2	4. a) Explain the graphical representation (interpretation) of IR spectra of	
)		benzoic acid.	
9	b	Mention the number of signals and multiplicity of the signals in the	(4+2)
		NMR spectrum of 1, 1, 2 trichloroethane.	(412)
2	5. a	a) Give the synthesis of ibuprofen.	9.9
9	t	o) Give any two principles of green chemistry.	(4+2)